



Shri Vaishnav Vidyapeeth Vishwavidyalaya, Indore

Degree Program B.Sc.

Semester-I (B.Sc.)

SUBJECT CODE	Category	SUBJECT NAME	TEACHING & EVALUATION SCHEME								
			THEORY			PRACTICAL		Th	T	P	CREDITS
			END SEM University Exam	Two Term Exam	Teachers Assessment	END SEM University Exam	Teachers Assessment				
BSPHU 101	DC	Foundation Course-I	60	20	20	00	00	3	1	0	4

Legends: L - Lecture; T - Tutorial/Teacher Guided Student Activity; P - Practical; C - Credit; Q/A - Quiz / Assignment / Attendance, MST Mid Sem Test.

Course Objectives:-

1. To develop the comprehensive understanding of English related to communication and behavior and ability to apply them for laying the foundation for research and development.
2. To work ethically as member as well as leader in a diverse team.

Course Outcomes:-

1. Student will be able to understand and solve the problems related to communication .
2. Student will be able to understand and solve the problems related to their behavior.

Foundation Course

Unit 1: Grammar: Basic Language Skills : Grammar and usage- Types of Sentences, Phrases & Clauses, Parts of Speech , Direct - Indirect, Active - Passive voice, S-V Agreement, Tenses.

Unit 2: Poetry : Robert Frost – The Road Not Taken, Nissim Ezekiel- Night of the Scorpion.

Unit 3: Short Stories : Images – A Handbook of Stories Ed. M. M. Lukose, Macmillan)

- i) R. K. Narayan- The Axe
- ii) The Mark Twain –Luck

Unit 4: Indian Culture: Indian Architecture, Sculpture, Music and Dance Forms.

Unit 5: Creative Writing: Precis Writing- General Rules and Steps, Paragraph Writing on Idiomatic Expressions.

Recommended Readings:

- A.J. Thomson and A.V. Martinet.(1991) **A Practical English Grammar** (4th ed). Newyork: Oxford IBH Pub.
- Sanjay Kumarm Pushp Lata.(**English for Effective Communication**.New Delhi: Oxford University Press India.
- Kalpana Rajaram, R. Vidya.(2012)**Facets of Indian Culture**. Spectrum Books Pvt. Ltd.,
- Ashraf Rizvi.(2005).**Effective Technical Communication**. New Delhi:Tata Mc Graw Hill.


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PAPERS

1. BS.P.LS.102 Cell Biology & Basic Microbiology

SUBJECT CODE	Category	SUBJECT NAME	TEACHING & EVALUATION SCHEME								
			THEORY			PRACTICAL		Th	T	P	CREDITS
			University Exam	Two Term Exam	Teachers Assessment*	University Exam	Teachers Assessment*				
BS.P.LS. 102	DC	Cell Biology & Basic Microbiology	60	20	20	30	20	4	1	2	7

2. Legends: L - Lecture; T - Tutorial/Teacher Guided Student Activity; P - Practical; C - Credit;

3. *Teacher Assessment shall be based following components: Quiz/Assignment/Project/Participation in Class, given that no component shall exceed more than 10 marks.

Course Objectives:

1. To give a comprehensive idea about the structural aspects of plant and animal cells and cellular organelles
2. To give a comprehensive idea about the important classes of microorganisms and their importance

Course Outcomes:

1. Student will be able to understand the structure and function of cell organelles, cell divisions and cell cycle
2. Student will be able to understand the salient features of microorganisms, their importance in nature and their control

A. Cell Biology

Unit-1

Discovery of cell and cell theory. Structure of prokaryotic and eukaryotic cells. Cell division and cell cycle. Cell synchrony and Cell signaling.

Structure and function of Cell wall and Plasma membrane, Diffusion and Osmosis; ion channels and ion pumps.


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Unit - II

Structure and function of - Mitochondria and Chloroplast.
Nucleus and Nucleolus. Structure of chromosomes, Polytene and Lampbrush Chromosomes

Unit - III

Structure and function of - Endoplasmic reticulum, Golgi apparatus, Lysosomes, Ribosomes, Microtubule, Microfilaments and Intermediate filaments.

B. Basic Microbiology

Unit- IV

History of microbiology; Important discoveries in Microbiology; Structure of bacteria; Classification of microorganisms and major groups of bacteria; General characteristics of viruses

Unit - V

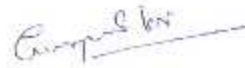
Nutritional classes of bacteria; Types of media and cultivation of bacteria; Factors affecting growth; Batch, continuous and synchronous culture; Measurement of bacterial growth; Growth curve and phases of growth cycle; Generation time and growth rate

Control of microorganisms - Physical methods [temperature, filtration, radiation]; Chemical methods for disinfection and sanitation

Role of microorganisms in nitrogen carbon, sulphur and phosphorus cycle

BS.P.LS.102 Life Science -1 List of Practicals

1. Aseptic transfer techniques for microorganisms.
2. Morphological study of microorganisms by;
 - a) Wet mount
 - b) Monochrome staining
 - c) Gram staining
 - d) Spore staining
 - e) Fungal staining
3. Preparation of liquid and solid media for culturing bacteria/fungi.
4. Isolation of bacteria from air/water/soil by Streak Plate Technique.
5. Techniques for enumeration of microorganisms;
 - a) Pour plate method
 - b) Spread plate method
 - c) Cell count by Neubauer's Chamber
 - d) Plaque Count
 - e) Turbidometric method

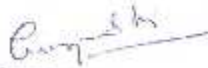

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6. Observation of motility of microorganisms by:
 - a) Hanging drop technique
 - b) Swarming growth
7. Growth curve of bacteria and calculation of generation time.
8. Effect of UV light/pH/temperature/salt concentration on bacterial growth.
9. Observation of various stages of mitosis and meiosis
10. Identification of the cell membrane, nucleus and cytoplasm in an animal cell and the cell wall, cell membrane, sap vacuole, cytoplasm, nucleus and chloroplasts in a plant cell
11. Methods of cell lysis: osmotic/enzymatic/chemical.
12. Permanent slides of: Prokaryote, Eukaryote, Muscle, Nerve, Stomata.

Books

1. Microbiology - Pelczar, Chan and Kreig
2. Microbiology - an introduction - Tortora, G. J., Funk, B. R. and Case C. L.
3. General Microbiology - Stainer R Y, Ingraham J L, Wheelis M L. and Painter P R
4. Biology of Microorganisms - Brock and Madigan
5. Elementary Microbiology - H. A. Modi
6. Text Book of Microbiology - Dubey and Maheshwari
7. Microbiology - A practical approach - Patel and Phanse
8. Microbiology - S. S. Purohit
9. General Microbiology - Powar and Dagainawala
10. Cell Biology, Powar C.B. Himalaya Publishers, Students Edition
11. 6. Cell Biology, Rastogi, S.C. (Edn. 3) New Age International, 2007.
12. 7. Essential Cell Biology (2nd Ed) by B. Alberts et al, Taylor & Francis Group; 2 edition.


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			END SEM University Exam	Two Term Exam	Teachers Assessment*	END SEM University Exam	Teachers Assessment*				
BSPCH 103	DC	Physical Chemistry I	60	20	20	30	20	3	1	4	6

Legends: L - Lecture; T - Tutorial/Teacher Guided Student Activity; P - Practical; C - Credit;

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Physical Chemistry- I

Course Objectives:-

To give basic knowledge of state of matter.

To understand and apply the knowledge of equilibria.

Course Outcomes:-

After completion of this course the students are expected to be able to demonstrate following knowledge, skills and attitudes. The student will demonstrate capability of

CO1. Theoretical understanding of various state of matter.

CO2. Became aware of the importance of equilibria and its laws in the field of chemistry and dealing with its numerical approach.

Unit I: Gaseous state:

Kinetic molecular model of a gas; postulates and derivation of the kinetic gas equation; collision frequency; collision diameter; mean free path and viscosity of gases, including their temperature and pressure dependence, relation between mean free path and coefficient of viscosity, calculation of σ from η ; variation of viscosity with temperature and pressure.

Maxwell distribution and its use in evaluating molecular velocities (average, root mean square and most probable) and average kinetic energy, law of equipartition of energy, degrees of freedom and molecular basis of heat capacities.

Behavior of real gases : Deviations from ideal gas behavior, compressibility factor, Z, and its variation with pressure for different gases. Causes of deviation from ideal behaviour. vander Waals equation of state, its derivation and application in explaining real gas behaviour.


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Unit II: Liquid state:

Qualitative treatment of the structure of the liquid state; Radial distribution function; physical properties of liquids; vapour pressure, surface tension and coefficient of viscosity, and their determination. Effect of addition of various solutes on surface tension and viscosity. Explanation of cleansing action of detergents. Temperature variation of viscosity of liquids and comparison with that of gases. Qualitative discussion of structure of water.

Unit III: Solid state:

Nature of the solid state, law of constancy of interfacial angles, law of rational indices, Miller indices, elementary ideas of symmetry, symmetry elements and symmetry operations, qualitative idea of point and space groups, seven crystal systems and fourteen Bravais lattices; X-ray diffraction, Bragg's law, a simple account of rotating crystal method and powder pattern method. Analysis of powder diffraction patterns of NaCl, CsCl and KCl. Defects in crystals. Glasses and liquid crystals.

Unit IV: Ionic equilibria:

Strong, moderate and weak electrolytes, degree of ionization, factors affecting degree of ionization, ionization constant and ionic product of water. Ionization of weak acids and bases, pH scale, common ion effect; dissociation constants of mono-, di- and tri-protic acids (exact treatment). Salt hydrolysis-calculation of hydrolysis constant, degree of hydrolysis and pH for different salts. Buffer solutions; derivation of Henderson equation and its applications; buffer capacity, buffer range, buffer action and applications of buffers in analytical chemistry and biochemical processes in the human body. Solubility and solubility product of sparingly soluble salts – applications of solubility product principle.

Unit V: Phase equilibria:

Definitions of phase, component and degrees of freedom. Phase rule and its derivations. Definition of phase diagram. Phase equilibria for one component system – water, CO₂. First order phase transition and Clapeyron equation; Clausius-Clapeyron equation - derivation and use. Solid-liquid phase diagram. Eutectic mixture.

Recommended Texts:

1. Atkins, P.W. & Paula, J. de Atkin's Physical Chemistry 8th Ed., Oxford University Press (2006).
2. Ball, D. W. Physical Chemistry Thomson Press, India (2007).
3. Castellan, G. W. Physical Chemistry 4th Ed. Narosa (2004).
4. Mortimer, R. G. Physical Chemistry 3rd Ed. Elsevier; NOIDA, UP (2009).


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4. BS.P.BT.104. Biochemistry and Instrumentation

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			THEORY			PRACTICAL		Th	T	P	CREDITS
			University Exam	Two Term Exam	Teachers Assessment*	University Exam	Teachers Assessment**				
BS.P.BT.104	DC	Biochemistry and instrumentation	60	20	20	30	20	4	1	4	7

5. Legends: L – Lecture; T – Tutorial/Teacher Guided Student Activity; P – Practical; C – Credit;
 6. *Teacher Assessment shall be based following components: Quiz/Assignment/Project/Participation in Class, given that no component shall exceed more than 10 marks.

Course Objectives:

1. To have the knowledge of chemical nature of important biomolecules
2. To know the principles of techniques and instruments used in biological laboratories

Course Outcome:

1. Students will be able to understand the chemical nature of biomolecules and their physical and chemical properties
2. Students will be familiar with the analytical techniques and the working principles of the instruments used in biological laboratories

A. Biochemistry

Unit-I

Structure, classification and function of Carbohydrates and Lipids
 Structure and types of DNA and RNA

Unit - II

Structure, classification and function of Amino acids and Proteins
 Structure and function of vitamins

Unit-III

Enzymes - Classification, Energy of activation and Enzyme kinetics, Michaelis - Menten equation,

Cofactors and coenzymes; Isozyme, ribozyme and abzymes

Allosteric enzymes


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B. Instrumentation

Unit - IV

Light Microscopy (bright field, dark field), Phase contrast and Fluorescence.
Electron Microscopy (SEM and TEM)

Spectroscopy: Beer Lambert's Law, Colorimeters, UV and Visible spectroscopy, Flame photometer.

Unit - V

Chromatography: Paper, Thin layer, Ion exchange, affinity and Gel filtration

Electrophoresis: Agarose gel, SDS PAGE and Native PAGE.

Centrifugation - Basic principles, preparative and analytical centrifuges

Radioactivity - autoradiography, Geiger counter and Scintillation Counter.

BS.P.BT.104. Biochemistry and Instrumentation List of Practicals

1. Qualitative and Quantitative [Nelson Somogyii's/DNS method] estimation of carbohydrates.
2. Qualitative and Quantitative [Folin Lowry's method] estimation of Proteins.
3. Determination of the enzyme activity by colorimetric methods
4. Effect of temperature on the activity of the given enzyme.
5. Effect of pH on the activity of the given enzyme
6. Effect of enzyme concentration on the activity of the given enzyme.
7. Effect of substrate concentration on the activity of the given enzyme and determination of V_{max} and K_m .
8. Determination of concentration of DNA by DPA method.
9. Determination of concentration of RNA by Orcinol method.
10. Separation of leaf pigments by paper chromatography.
11. Separation of aliphatic, aromatic and polar amino acids by TLC.
12. Isolation of biomolecules from natural sources.
13. Agarose gel electrophoresis of DNA/RNA sample.
14. Separation of proteins on the basis of size by SDS-PAGE.


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BOOKS

1. Principles of Biochemistry. Lehniger, 3rd edition by Nelson and Cox (Worth) 2000
2. Biochemistry Stryer 5th edition W.H. Freeman 2001.
3. Harper's Biochemistry, 1999 (McGraw-Hill)
4. Zubey GL, Parson WW and Vance DE (1994) Principles of Biochemistry – W.M.C. Brown Publishers, Oxford, England
- 5.
8. Fundamentals of Biochemistry, Jain, J.L.,
9. Biochemical Methods of Analysis: *Theory and Applications*, Saroj Dua, S., Garg, N. Narosa Publishing House
10. Biochemistry, Sharma, D.K. Narosa Publishing House
11. Experiments in Biotechnology Nighojkar and Nighojkar


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